

Article

Phosphogypsum and Black Steel Slag as Additives for Ecological Bentonite-Based Materials: Microstructure and Characterization

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Abstract: The Black Steel slag (Ss) and phosphogypsum (PG) are industrial wastes produced in Morocco. In order to reduce these two wastes and to evaluate their pozzolanic reactivity in the presence of water, they were incorporated into bentonite (B) mixed with lime (L). The studied mixtures (BLW, BL-PG-W and BL-PG-Ss-W) were analyzed by X-ray diffraction, Infrared spectroscopy, Raman spectroscopy and SEM/EDX analysis. Compressive strength tests were performed on hardened specimens. The results obtained show that the hydration kinetics of the B-L-W and B-L-PG-W mixtures are slow. The addition of PG to a bentonite—lime mixture induces the formation of new microstructures such as hydrated calcium silicate (C-S-H) and ettringite, which increases the compressive strength of the cementitious specimens. The addition of the Ss to a mixture composed of 8%PG and 8%L-B accelerates the kinetics of hydration and activates the pozzolanic reaction. The presence of C₂S in the slag helps to increase the mechanical strength of the mixture B-L-PG-Ss. The compressive strength of the mixtures BL-W, BL-PG-W and BL-PG-Ss-W increases from 15 to 28 days of setting. After 28 days of setting, 8% of Ss added to the mixture 8% PG-8%L-B is responsible for an increase of the compressive strength to 0.6 MPa.

Keywords: bentonite; phosphogypsum; slag; lime; hydration; kinetics; Pozzolanic reactions

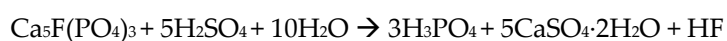
1. Introduction

Recycling industrial waste as an alternative to reduce natural resource consumption is a recently-developed solution to some environmental and economic problems. China, Europe, Japan and the United States are the worldwide leaders of steel industry production according to the statistics reported by the World Steel Association [1].

In Morocco, the steel industry in the Jorf Lasfar (El Jadida, NE Morocco), represented by the SONASID Company (El Jadida, NE Morocco), i.e., a subsidiary of ArcelorMittal, produces more than 150,000 tons of black steel slag (Ss) per year and recovers 90% of the slag produced [2]. Black steel slag presents major environmental risks [3]. It contains heavy metals like As and Cr that can be released into water sources [4]. In particular, arsenic, in small concentrations, is toxic and highly mobile in various pH conditions, which may leach out into the environment [5–7]. Its chemical composition consists mainly of CaO, MgO, SiO₂, MnO and FeO [8,9]. The slag has a porous

morphology and alkaline properties. Therefore, it is largely used for the retention of phosphate ions from waste-contaminated water [10–12]. The presence of calcium silicate (Ca_2SiO_4), Tricalcium silicate (Ca_3SiO_5), tetracalcium aluminoferrite ($\text{Ca}_4\text{Al}_2\text{Fe}_2\text{O}_{10}$) and other mineral components provide to the steel slag hydraulic setting properties [13,14], allowing an increase of compressive strength of dredged marine clays [15]. Steel slag is also successful used as an aggregate in the cement industry [16]. It helps to delay the hydration mechanism of concrete and reduces its mechanical strength [17]. Likely slag activates the hydration of C_2S and C_3S [18]. In addition, the occurrence of Fe_2O_3 in the steel slag decreases the expansion of clay-based material and increases its mechanical strength [19]. In particular, the use of steel slag as aggregate for metakaolinitic–cement enhances its mechanical properties [20]. In addition, steel slag is used in other fields such as civil engineering, mainly to stabilize the swelling behavior of clays in road infrastructure [21].

Phosphogypsum (PG) essentially consists of calcium sulfate dihydrate ($\text{CaSO}_4 \cdot 2\text{H}_2\text{O}$). It is a by-product resulting from the attack of sulfuric acid on phosphate rock to produce phosphoric acid, mainly used in the fertilizer industry. The process of PG production is described by the following reaction [22]:



The world phosphate industry produces about 250 Mt/yr of PG [22]. In Morocco, it exceeds 25 Mt/year. PG is produced in large quantities, five tons of PG for one ton of phosphoric acid [23,24]. PG also contains phosphates under the form of H_3PO_4 , $\text{Ca}(\text{H}_2\text{PO}_4)_2 \cdot \text{H}_2\text{O}$ and $\text{CaHPO}_4 \cdot 2\text{H}_2\text{O}$, $\text{Ca}_3(\text{PO}_4)_2$, residual acids, fluorids (NaF , CaF_2 , Na_3AlF_6 , Na_3FeF_6 and Na_2SiF_6), trace metals (Cd , Zn , Cu , Cr) and organic matter [25]. It can also contain nuclides such as Ra, U and Th, but do not pose any major leaching environmental risk, except for Cd [22,26].

Recently, several studies aimed at valorizing PG to improve road pavements and stabilize the swelling behavior of soils have been realized [27–32]. PG was also used as aggregates for bricks masonry materials [33,34] and for the production of cement [35,36]. Romero–Hermida et al., [37] studied rheological characteristics of a lime putty obtained as a by-product from the treatment of phosphogypsum waste with a solution of $\text{Na}(\text{OH})$. The putties studied present a similar rheological response as a commercial one. Gu and Chen [38] attested that the addition of PG and fly ash increased the compressive strength of the loess-PG-fly ash mixture due to the formation of nanosilicates gel (C–S–H) gel and ettringite (AF-t). Furthermore, PG promotes the carbonization rate of steel slag [39].

Bentonite, mainly composed of calcium and sodium montmorillonite, is among the most used clays in civil engineering due to its high adsorption capacity [40]. At room temperature, at high pH, and in the presence of water and lime (CaO), bentonite undergoes geopolymerization like a Portland cement through complex pozzolanic reactions.

Our study aims to determine the influence of the addition of black steel slag and phosphogypsum on the mechanical performance of bentonite-based mortar with the presence of lime. This study will also strive to understand the different reaction mechanisms and to identify the hydrated microstructures after the hardening process using several analytical complementary techniques (XRD, IR, RAMAN and SEM/EDX).

2. Material and Methods

2.1. Raw Materials

The raw bentonite was sampled from the Trebia deposit in the Nador area (North–East Morocco). The traditional lime marketed in Oujda (eastern Morocco) is used to stabilize the swelling behavior of bentonite. The black Ss produced by the ArcelorMittal SONASID company in the Jorf Lasfar (Eljadida, Morocco), is obtained after melting iron scrap in an electric arc furnace. Phosphogypsum (PG), produced by OCP at Jorf Lasfar (Eljadida, Morocco).

2.2. Experimental Procedure

The collected samples were ground and sieved to a diameter of less than 250 μm . Three mortar specimens were prepared with Ss, PG, L and B according to the proportions reported in Table 1. The prepared specimens are cylindrical with 3.4 cm in diameter and 6.8 cm in height according to ASTM Standard D1621. The test tubes contain 80 g of the solid mixed with 38.4 mL of the distilled water having a W/S ratio of 46% [31]. The plasticity index and Particle-Size Distribution analysis are determined according to the ASTM D4318 and ASTM D6913, accordingly [41].

The mechanical resistance of the specimens was monitored after 15 and 28 days of setting. For the ES3, EP3 and EL3 test pieces, the mineralogical composition was determined by XRD after 3, 7, 15 and 28 days of setting. All test tubes were immersed in distilled water for 4 h and the pH of the solution was measured.

Table 1. Formulations based on the different proportions of raw materials studied.

| Mixtures | Ss (wt.%) | PG (wt.%) | L (wt.%) | B (wt.%) |
|----------|-----------|-----------|----------|----------|
| EL1 | 0 | 0 | 2 | 98 |
| EL2 | 0 | 0 | 5 | 95 |
| EL3 | 0 | 0 | 8 | 92 |
| EL4 | 0 | 0 | 11 | 89 |
| EL5 | 0 | 0 | 14 | 86 |
| EL6 | 0 | 0 | 20 | 80 |
| EP1 | 0 | 2 | 8 | 90 |
| EP2 | 0 | 5 | 8 | 87 |
| EP3 | 0 | 8 | 8 | 84 |
| EP4 | 0 | 11 | 8 | 81 |
| EP5 | 0 | 14 | 8 | 78 |
| EP6 | 0 | 20 | 8 | 72 |
| ES1 | 2 | 8 | 8 | 82 |
| ES2 | 5 | 8 | 8 | 77 |
| ES3 | 8 | 8 | 8 | 74 |
| ES4 | 11 | 8 | 8 | 73 |
| ES5 | 14 | 8 | 8 | 70 |
| ES6 | 20 | 8 | 8 | 64 |

2.3. Characterization Techniques

The X-ray diffraction was carried out by a Shimadzu XRD 6100 diffractometer equipped with a Cu X-ray tube (Faculty of sciences, Oujda, Morocco), operating at 40 kV and 30 mA, in the range 2° – 70° 2θ . The infrared analysis was obtained using a Fourier transform spectrometer (FT/IR-4700, JASCO, Faculty of sciences, Oujda, Morocco), ATR mode (attenuated total reflectance), equipped with a DLaTGS detector and Peltier temperature control (Faculty of sciences, Oujda, Morocco). Rapid 10 Hz scan with a resolution of 0.4 cm^{-1} have acquired in a wavelength range of 400 – 4000 cm^{-1} . Raman and Infrared (FTIR) analysis was performed at the Spectrometry Laboratory (CNRST, Rabat, Morocco), using the Vertex 70 Bruker apparatus equipped with an MCT detector with a resolution of 0.5 cm^{-1} . The spectral range from 4000 to 70 cm^{-1} . The pH was measured by a 920 Benchtop precision pH meter (Faculty of sciences, Oujda, Morocco). The Scanning Electron Microscopy (SEM) was carried out at the Scientific Research National Center (CNRST, Rabat, Morocco), with a resolution of 3.5 nm coupled with an EDX microanalysis of 133 eV resolution. The TESTWELL device with a speed of 1 mm/min obtained the compressive strength analysis. The thermal analysis was carried out using thermogravimetry (TGA) and differential scanning calorimetry (DSC) by heating the samples from 20 to $1000\text{ }^\circ\text{C}$ at a uniform temperature ratio of $10\text{ }^\circ\text{C}/\text{min}$.

3. Results and Discussion

3.1. Characterization of Raw Materials

The physical properties of raw bentonite are shown in Table 2. Bentonite studied is classified as fatty clays (CH) according to the Unified Soil Classification System (USCS). The particle size analysis by sedimentometry shows that this bentonite consists of 48.7% sandy fraction, 33.8% clay fraction

and 17.8% silt. The geotechnical analysis is obtained according to the ASTM D4318 and ASTM D6913. The liquid limit of bentonite is measured using the Casagrande apparatus, and the particle size distribution for the fraction less than 80 μm is obtained using the hydrometer method (Table 2).

The XRD analysis shows the occurrence of Ca-montmorillonite (Ca-Mt), Na-montmorillonite (Na-Mt), Na-feldspar (NaF), K-feldspar (KF) and Anorthite (Pl) (Figure 1). The chemicals composition of raw bentonite, phosphogypsum and black steel slag are presented in Table 3 [42,43]. The grains of bentonite show a “cornflake” or “oak leaf” morphology typical and representative of montmorillonite, associated with feldspars (Figure 2A). Lime XRD analysis reveals the presence of CaO, MgO and Ca (OH)₂ [31]. The steel slag consists of calcium ferrite (CaFe₂O₄), calcium dis-silicate (Ca₂SiO₄) srebrodskite (Ca₂Fe₂O₅), fayalite (Fe₂SiO₄), quartz (SiO₂), lime (CaO), dolomite (CaMg(CO₃)₂) and brownmillerite (Ca₂(Al, Fe)₂O₅) (Figure 1). The slag grains show pores with a diameter of 1.7–2.25 μm (Figure 2C). The Phosphogypsum used in this study contains gypsum (CaSO₄·2H₂O), bassanite (CaSO₄·1/2H₂O) and anhydrite (CaSO₄) [31]. The hazardous trace elements (Cd, Se, Cr etc.) containing in Moroccan PG and Ss leachates were studied, respectively, by Sebbar et al. [44] and Hassoune et al. [45] (Table 4).

Table 2. Physical properties of raw bentonite.

| Clay | Liquid Limit % | Plasticity Limit % | Plasticity Index (PI) | Clay (<0.005 mm) % | Silt (0.005–0.075 mm) % | Sand (0.075–2 mm) % | pH ₀ |
|-----------|----------------|--------------------|-----------------------|--------------------|-------------------------|---------------------|-----------------|
| Bentonite | 103 | 30.5 | 72.5 | 33.8 | 17.8 | 48.7 | 8.72 |

Table 3. Chemical analysis of raw materials.

| Compounds | Chemical Composition % in Mass | | | | | | | | | | | |
|-----------------------|--------------------------------|--------------------------------|--------------------------------|------|------|-----------------|------------------|------------------|------|-------------------|-------------------------------|--------------------------------|
| | SiO ₃ | Al ₂ O ₃ | Fe ₂ O ₃ | CaO | MgO | SO ₃ | K ₂ O | TiO ₂ | MnO | Na ₂ O | P ₂ O ₅ | Cr ₂ O ₃ |
| Bentonite | 59.9 | 24.4 | 3.00 | 1.11 | 2.57 | - | 0.69 | 0.20 | 0.01 | 2.20 | 0.03 | - |
| Phosphogypsum [42] | 9.50 | 2.80 | 0.90 | 30.5 | 0.30 | 42.5 | - | - | - | - | 0.50 | - |
| Black Steel slag [43] | 12.8 | 5.04 | 31.9 | 36.5 | 3.26 | 0.31 | - | 0.41 | 5.90 | - | 0.50 | 2.15 |

Table 4. Hazardous trace elements in phosphogypsum (PG) and steel slag (Ss) leachates.

| Leachate | Chemical Elements (ppm) | | | | | | | | |
|---------------------|-------------------------|--------|--------|--------|--------|-------|---------|----------|--------|
| | Zn | Ni | Fe | Cu | Cd | Pb | Se | As | Cr |
| Leachate of Ss [44] | <0.005 | <0.003 | <0.004 | <0.005 | <0.003 | <0.01 | <0.0015 | <0.00005 | <0.005 |
| Leachate of PG [45] | - | 0.125 | - | 0.496 | 1.558 | 0.021 | 0.017 | 0.074 | - |

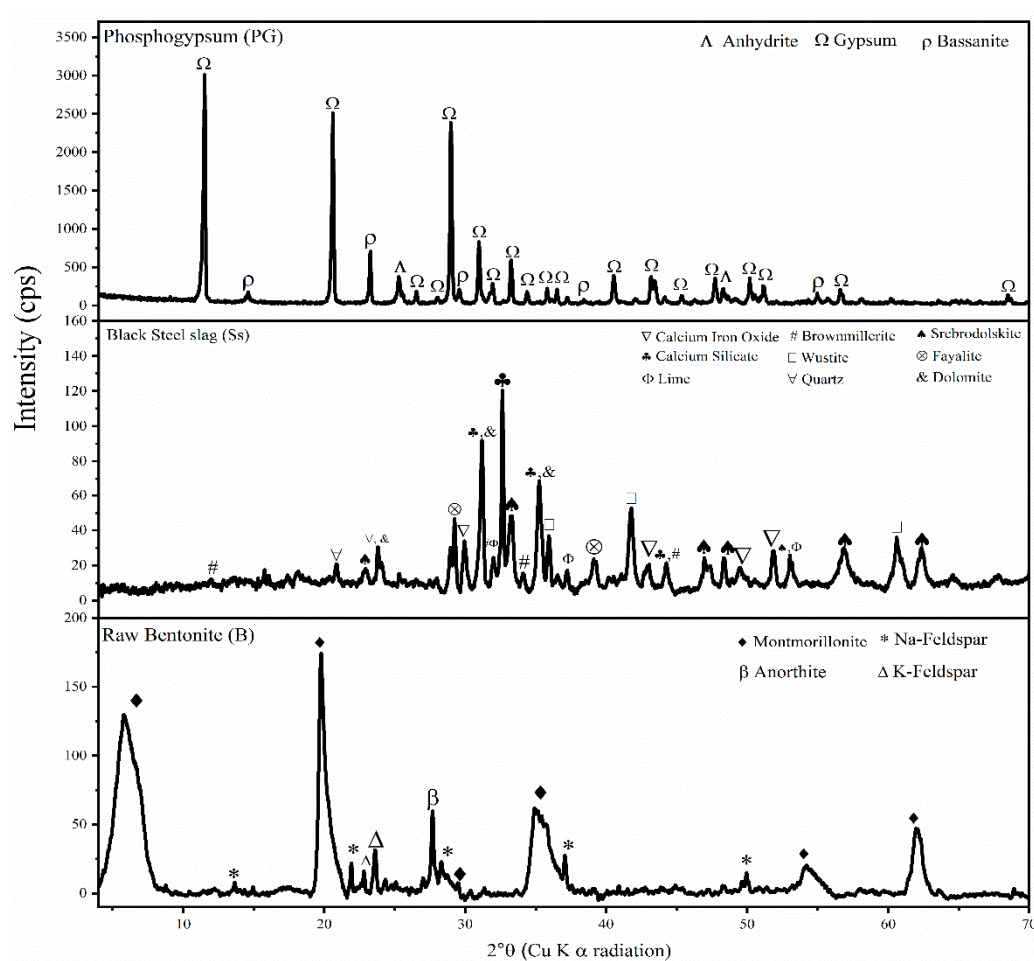


Figure 1. XRD patterns of phosphogypsum, steel slag, and raw bentonite.

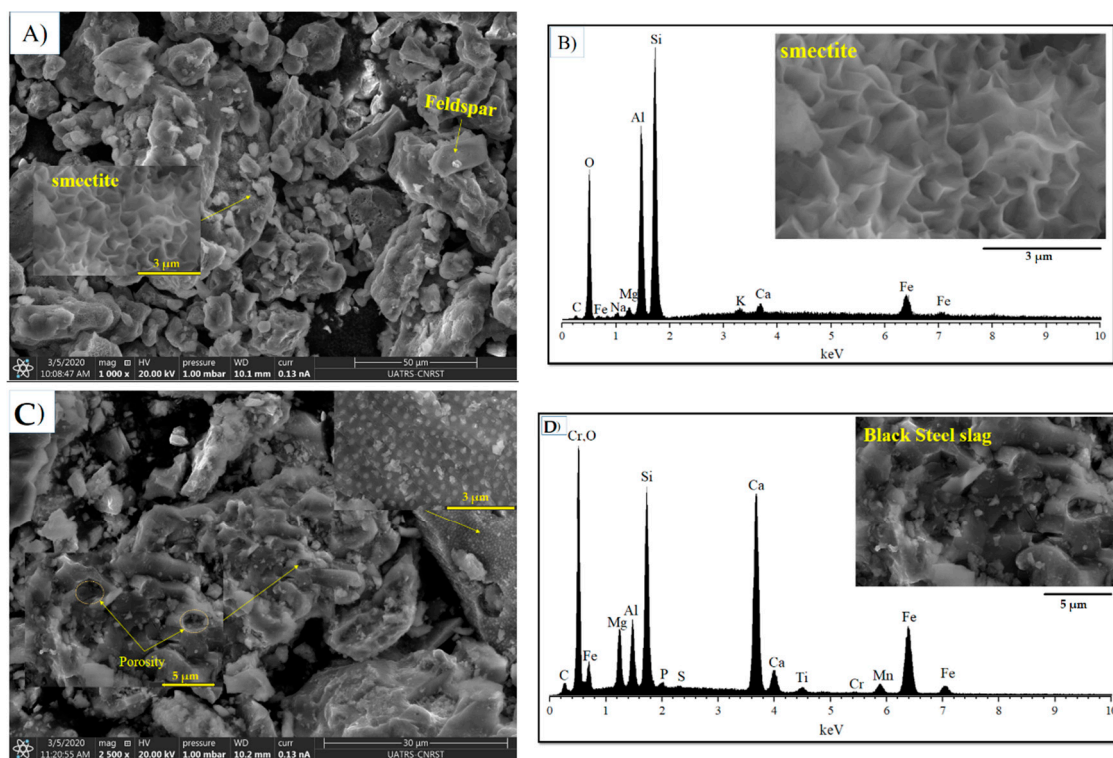


Figure 2. SEM and Energy-dispersive X-ray spectroscopy (EDX) images of raw material. (A,B) raw bentonite and (C,D) black steel slag.

3.2. Compressive Uniaxial Strength

Figure 3A compares the uniaxial compressive strengths of the bentonite–lime mixture specimens after 15 and 28 days of hydration. The histograms corresponding to the two periods of the setting are similar. The optimal value is 0.97 MPa for the EL mixture, corresponding to the maximum compressive strength, which is obtained when 8% of lime is added to the mixture (EL3). When increasing the amount of lime added, the compression strength increases. This result can be explained by the fact that the concentration of Ca^{2+} increases and the formation of C–S–H phase becomes more important due to the occurrence of pozzolanic reactions. From 15 to 28 days the amount of C–S–H increases, which is consistent with the increase in mechanical strength. Similar results are obtained by Manzano et al. [46]. In addition, they also showed that the longer the chains of the CSH silicate gel, the better the mechanical properties.

The addition of PG to the bentonite + 8% lime mixture increases the compressive uniaxial strength. The maximum value is obtained by adding 8% of PG (Figure 3B), which is in agreement with the results obtained by Oumnih and al. [30]. Using more than 8% of PG, there is a remarkable drop in mechanical strength. Excess PG increases the acidity of the mixture, making the chains of the C–S–H gel and the hydrates formed less stable, explaining the observed decrease in compressive strength. The sulfate released from PG is consumed by the reaction between the aluminate hydrated gel and calcium cation to produce ettringite, thereby increasing the mechanical properties of the specimens.

The addition of black steel slag to the bentonite + 8% lime + 8% PG mixture increases the uniaxial compressive strength of the specimen's test (Figure 3C). The optimum value is 1.91 MPa after 28 days of hydration for the ES3 mixture. It's obtained when adding 8% of black steel slag to the mixture B + 8%L + 8%PG. From 15 days to 28 days of hydration. The compression increases by 0.6 MPa probably due to the presence of hydraulic phases such as Ca_2SiO_4 , $\text{Ca}_2\text{Fe}_2\text{O}_5$ and Fe_2SiO_4 in the slag. These hydraulic phases may contribute to the setting of the cement paste. After the flocculation of the clay minerals, aggregates can form around the coarse grains. This releases the surface of the bentonite and allows its direct contact with the lime which activates the pozzolanic reactions. Similarly, slags, being alkaline, may increase intake by neutralizing the residual acidity of PG. The Figure 3D showed comparison between EL3 (B + 8%L), EP3 (B + 8%L + 8%PG) and ES3 (B + 8%L + 8%PG + 8%Ss).

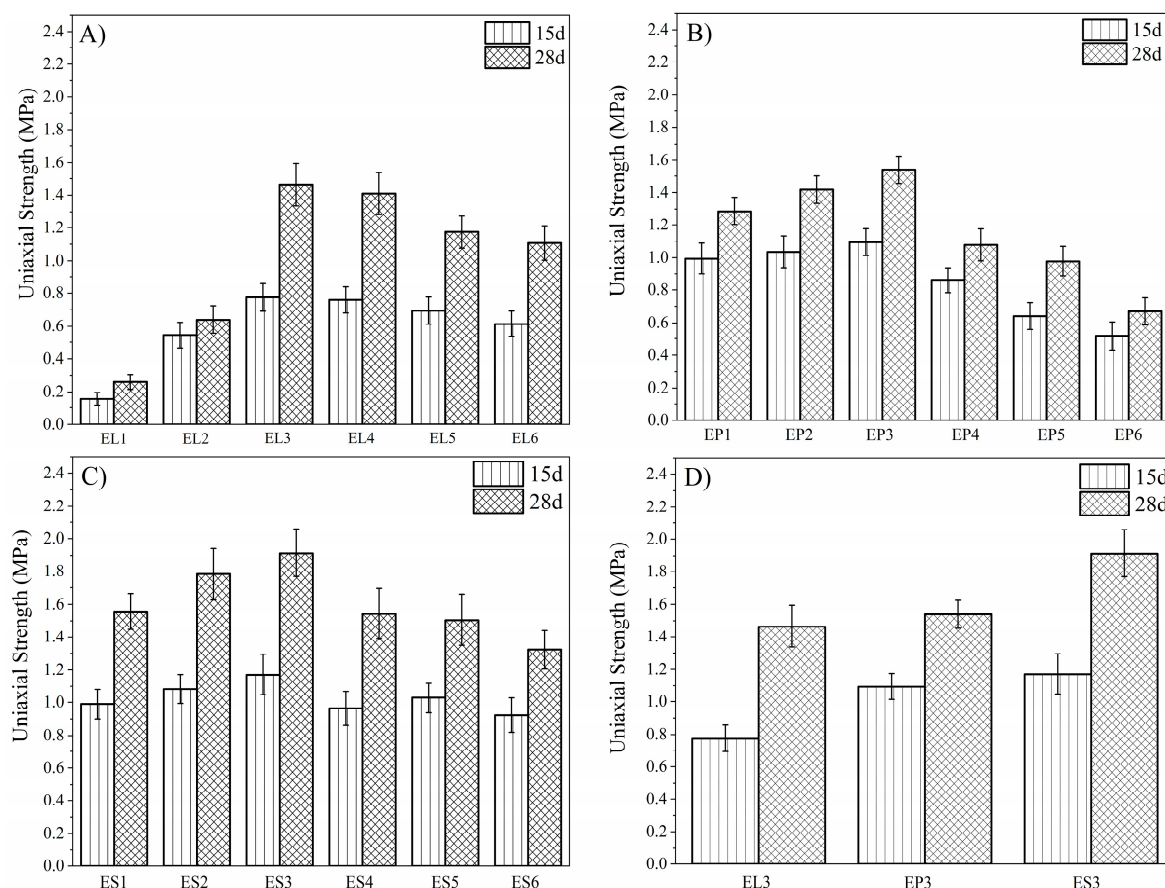


Figure 3. Compressive uniaxial strength of (A) bentonite + lime; (B) bentonite + lime + PG; (C) Bentonite + lime + PG + Ss after 15 and 28 days of hydration (D) Comparison between EL3 (B + 8%L), EP3 (B + 8%L + 8%PG) and ES3 (B + 8%L + 8%PG + 8%Ss).

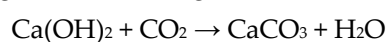
3.3. Kinetic Study by XRD

Figure 4A shows the evolution of the pH of the mixtures EL3, EP3 and ES3 over time. Figure 4B–D) indicates the evolution of the mineralogical composition of the mixtures EL3, EP3 and ES3.

The initial pH of the leaching solutions of the mixtures is high (pH = 12.5) due to the dissolution of calcium hydroxide ($\text{Ca}(\text{OH})_2$) in water. It is attested by the decrease of the intensities of the $\text{Ca}(\text{OH})_2$ peak ($d = 2.6 \text{ \AA}$) of the mixings EL3, EP3 and ES3.

The mixture EL3 (i.e., 92% bentonite and 8% of lime, soaked in water; Figure 4B), shows the presence of Ca-montmorillonite (Ca-Mt), Na-montmorillonite (Na-Mt), Na-feldspar (NaF) and K-feldspar (KF), lime (L), Anorthite (Pl), periclase (P), and calcium hydroxide (CH). In presence of water ($W/s = 46\%$), bentonite grains (Na-Mt and Ca-Mt) consolidate by coagulation aggregation and become denser [47]. The pH of the leach solution decreases slightly over time due to the consumption of OH^- ions released by the lime during pozzolanic reactions. The pH decreases from ~12 to 11.3 after 7 days of setting.

After seven days of hydration of the EL3 mixture, calcite (CaCO_3) is formed and showed by the appearance of the characteristic peak at $d = 3.03 \text{ \AA}$. The formation of calcite can be explained by the neo-carbonation of lime according to the following reaction:



After 28 days of hydration of the mixture EL3, the formation of hydrated calcium silicate (C–S–H) occurs at $d = 3.08 \text{ \AA}$, due to the alteration of sodium and calcium bentonite by Ca^{2+} and OH^- ions released by lime in intra-granular solutions. The formation of C–S–H consumes the OH^- ions, which causes a decrease in pH values.

Ettringite ($\text{Ca}_6\text{Al}_2(\text{SO}_4)_3(\text{OH})_{12}\cdot 26\text{H}_2\text{O}$) appears at $d = 5.61 \text{ \AA}$, is formed after 3 days of hydration of the mixtures EP3 and ES3 containing PG and PG + Ss, respectively (Figure 4B,C). Ettringite crystals may accelerate the hydration of C–S–H phases, increase their compressive strength, reduce the setting time, yield strength, and plastic viscosity of the mixtures [48].

After seven days of setting, the pH of the leachate solution of the EP3 specimen drops to 10.7 (Figure 4A) and the C–S–H phases appear (Figure 4B). The occurrence of C–S–H gel which consumes the hydroxide ions is accelerated by the presence of PG. The drop in pH could also be attributed to the release of residual acidity over time. After 28 days, the pH of the EP3 solution stabilizes at 9.48.

The addition of 8% of black steel slag to the mixture PG–bentonite–lime does not reflect any change in the XRD spectra of the EP3 mixture, and the evolution of the pH of the leached solutions gives the same variation as the EP3 mixture.

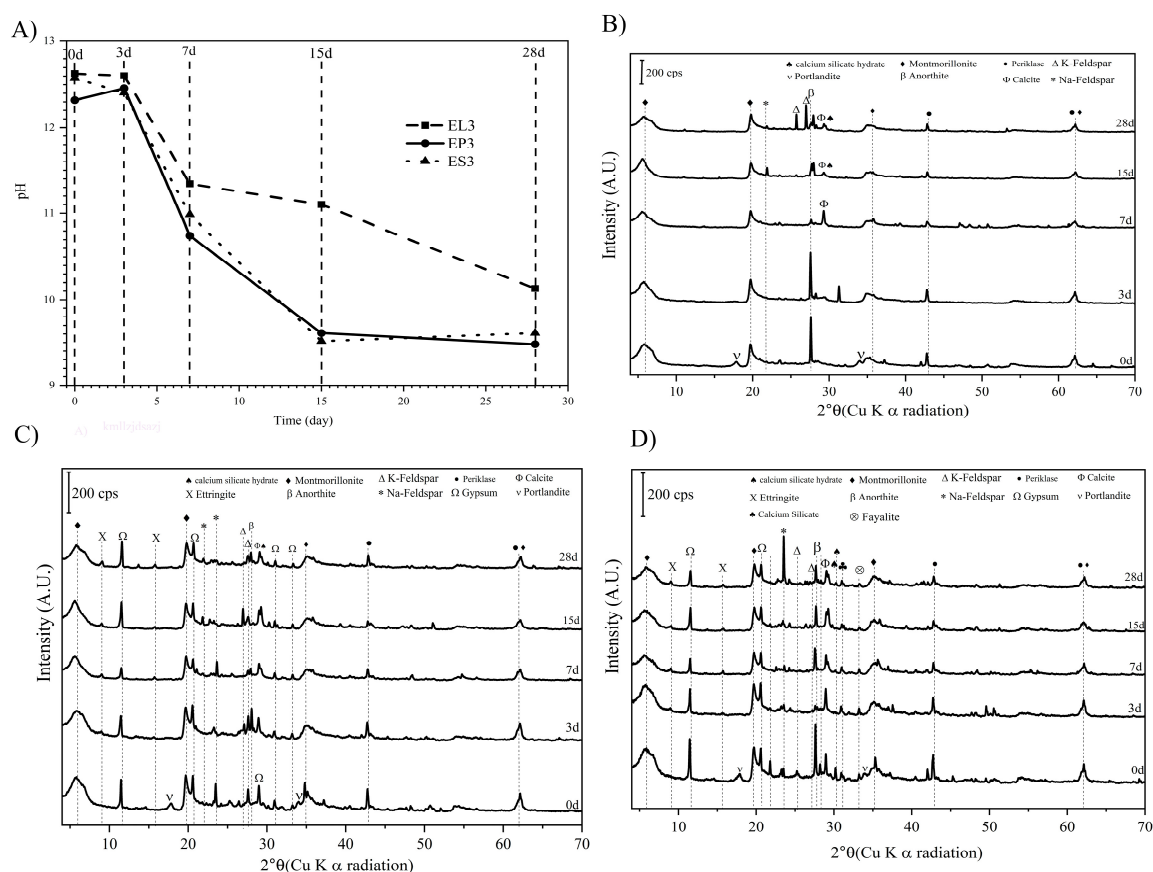


Figure 4. Evolution of pH of EL3, EP3 and ES3 mixtures (A); XRD spectra of EL3 (B), EP3 (C) and ES3 (D).

3.4. Infrared and Raman Results

Figure 5A shows the infrared spectra of raw bentonite and samples EL3, EP3 and ES3. The absorption band at $3700\text{--}3600 \text{ cm}^{-1}$ is due to stretching of the OH band of water adsorbed [20]. The two peaks observed at 1622 and 1539 cm^{-1} are related to the vibration of OH of water adsorbed in the interlayer space of Na-montmorillonite or Ca-montmorillonite and PG hydration water. The band at 1107 cm^{-1} is attributed to the bending vibrations of the Si–O band. The Al–Al–OH bending vibrations are observed at 918 cm^{-1} , the Al–O–Si deformation vibrations are at $\sim 520 \text{ cm}^{-1}$, and the Si–O–Si deformation vibrations are at 467 cm^{-1} [49].

When 8% of PG is added, a peak at 3411 cm^{-1} appears which is attributed to the symmetrical and asymmetric stretching of the OH bond of hydrated water in the gypsum molecule $\text{CaSO}_4\cdot 2\text{H}_2\text{O}$ [29]. Stretching and bending vibrations of SO_4 of PG are also observed at 1102 cm^{-1} and 669 cm^{-1} , respectively. The most intense band at 992 cm^{-1} is attributed to asymmetric Si–O–Si stretching and

the band at 518 cm^{-1} is due to the vibrational mode of Si–O–Si bending in calcium silicate (Ca_2SiO_4) found in steel slag [50].

Figure 5B shows the Raman spectra for raw bentonite and mixtures EL3, EP3 and ES3. The peak at 1086 cm^{-1} is due to Si–O* bond stretching vibration of $\text{Si}_2\text{O}_5^{2-}$ of the Si–O–Si (Q^4) bond in SiO_2 [51–53]. This peak disappears on the spectra EL3, EP3 and ES3 due to the destabilization of the siliceous structure by lime. C–S–H gel is formed in EL3, EP3 and ES3 mixtures change the environment of tetrahedra $[\text{SiO}_4]^{4-}$ and give the peak labeled Q^1 (Figure 5B). A peak at 702 cm^{-1} is observed for mixtures EL3, EP3 and ES3, attributed to the elongations of the Si–O* bond (Q^1) [54].

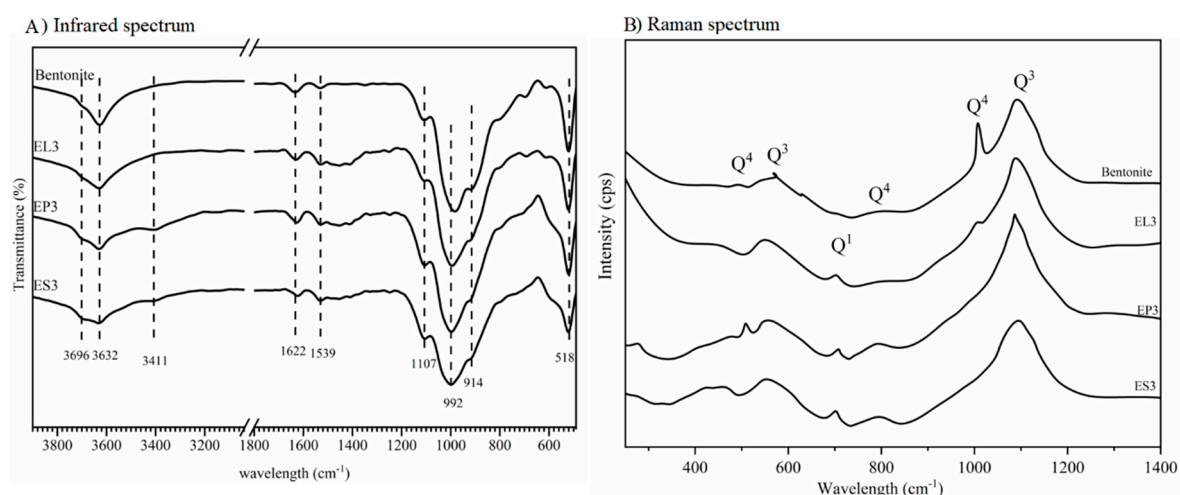


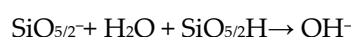
Figure 5. Infrared (A) and Raman (B) spectra of raw bentonite and EL3, EP3 and ES3 mixtures.

The mechanism of C–S–H chains formation involves three steps according to Monnin [55]:

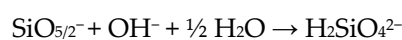
- Step 1: formation of Q^3 tetrahedra by an attack of hydroxyl ions on Q^4 tetrahedra:



Breaking a siloxane bond induces the formation of two Q^3 tetrahedra. Depending on the pH, the Q^3 sites are either ionized or protonated:



- Step 2: the dissolution of tetrahedra Q^3 in the form of tetrahedra Q^0 follows further cutting of siloxane bonds by hydroxyl ions:



- Step 3: the dissolved silica can diffuse into the interstitial solution and form with the alkalis and lime hydrated calcium silicates and/or hydrated calco-sodium silicates of type Q^1 :

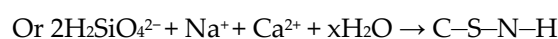
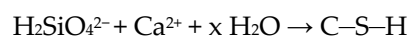


Figure 6 summarizes the arrangement of tetrahedra $[\text{SiO}_4]^{4-}$ in silicate structures and the corresponding peak type Q^n .

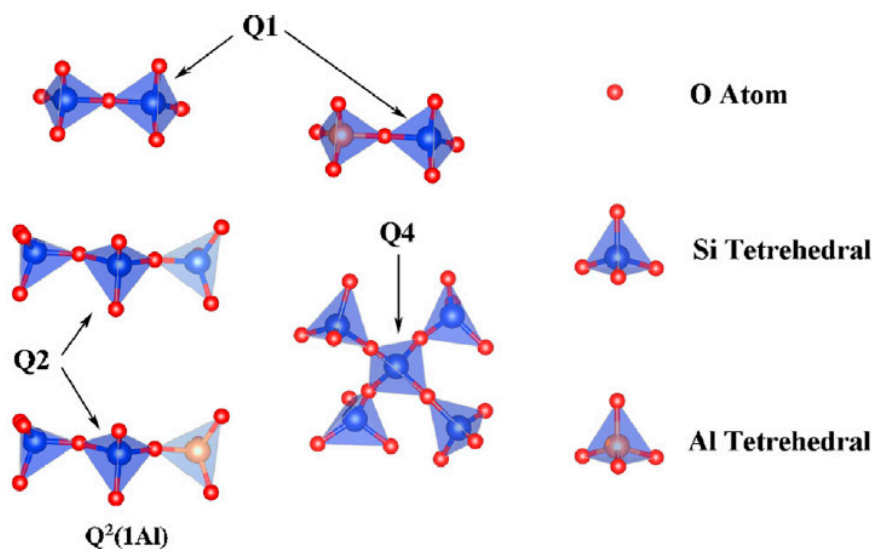


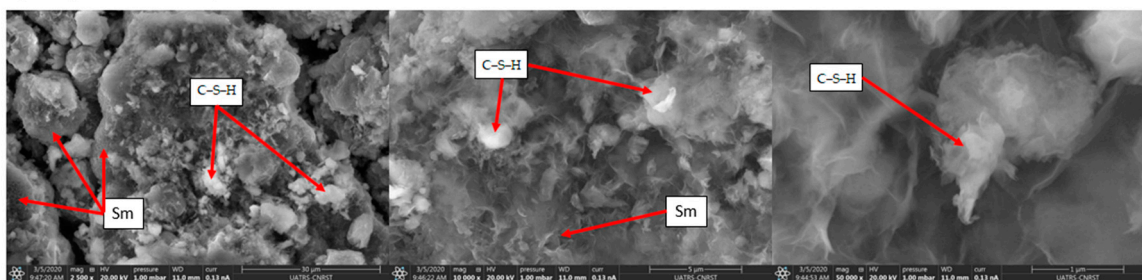
Figure 6. Schematic diagram of Q^n in aluminosilicate structure.

3.5. Microstructure

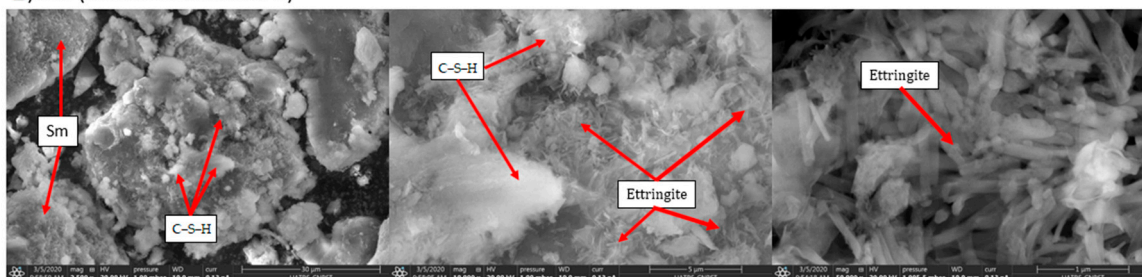
SEM observations of mixtures EL3 after 28 days of hydration (Figure 7A) shows the development of C–S–H gel around bentonite grains. The EP3 and ES3 mixtures rather display an elongated crystal of acicular shape corresponding to ettringite (Figure 7B,C). Ettringite is formed in the free spaces of the mixtures EP3 and ES3 and its presence decreases the porosity of the specimen block. Consequently, ettringite contributes to improving the cohesion between the particles of clays and therefore increases the mechanical strength.

The particles of the black steel slag have modified the surface of the bodies (Figure 7C), allowing ettringite and C–S–H gel neoformation. The influence of chemical alteration may be attested by the presence of calcium di-silicate (Ca_2SiO_4 , belite) in steel slag. Belite, a major component of portland cement that reacts with water to give C–S–H gel. The Figure 7D showed the Energy-dispersive X-ray spectroscopy (EDX) of EL3, EP3 and ES3 mixtures.

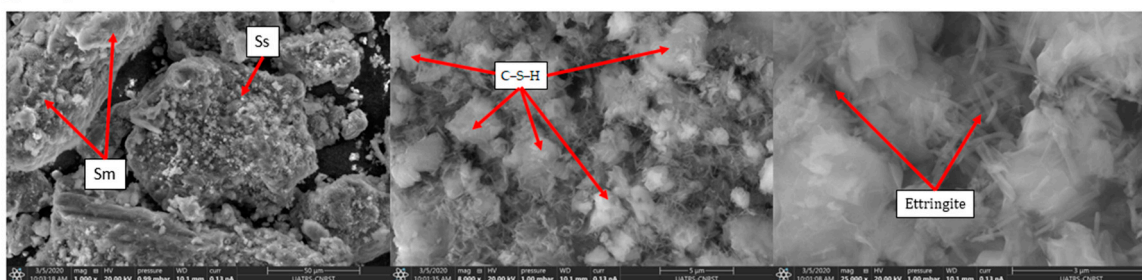
A) EL3 (Bentonite+8%L)



B) EP3 (Bentonite+8%L+8%PG)



C) ES3 (Bentonite+8%L+8%PG+8%Ss)



D) EDX microanalysis spectra

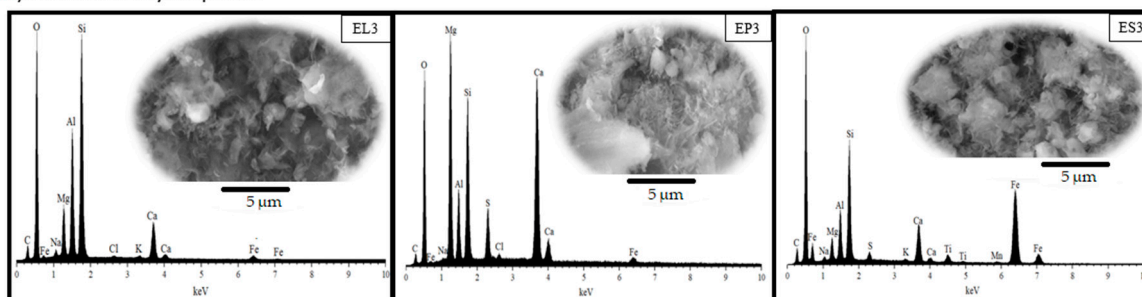
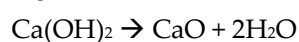


Figure 7. SEM images of (A) EL3, (B) EP3, (C) ES3 mixtures and (D) EDX spectra of EL3, EP3 and ES3.

3.6. DSC Analysis

As shown in Figure 8, the DSC analysis of EL3 mixture showed four endothermic peaks associated with a loss of mass. The first weight loss that occurs between 85 and 156 °C, is attributed to the water molecules adsorbed to the Ca–Mt or Na–Mt.

The addition of 8% of PG in EP3 and ES3 mixtures shows two small endothermic peaks at 128 °C and 152 °C, which the latter corresponds to a loss of water molecules from ettringite crystal in the samples [56]. The endothermic peak at 470–545 °C can be attributed to the decomposition of calcium hydroxide to the lime with the following reaction:



The fourth weight loss occurs at 650–800 °C, which corresponds to the decomposition of calcite [56,57].

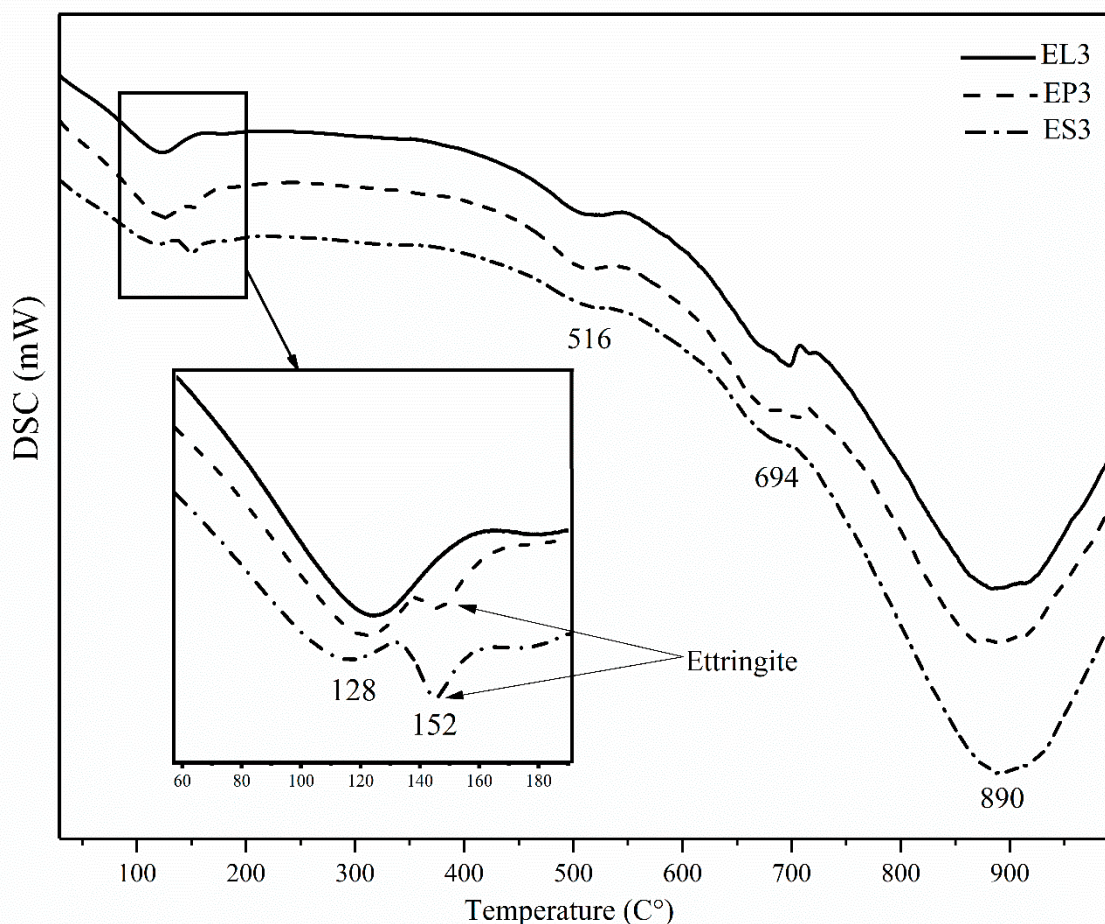


Figure 8. Differential scanning calorimetry (DSC) curves of EL3, EP3 and ES3 mixtures after 28 days of hydration.

4. Conclusions

In this paper, the industrial waste produced in Morocco, namely black steel slag (Ss) and phosphogypse (PG), was recovered by mixing it with bentonite (B) and lime (L) without any prior chemical treatment. The use of PG and black steel slag as additive raw materials in bentonite–lime mixture is valuable for environment-protecting and resource conservation. These wastes increase the mechanical properties of bentonite–lime mixtures. This enhancement promises to bring about new opportunities for using these wastes as an additive in concrete cement for building and road embankment construction.

Microstructures and characterization of bentonite–lime, bentonite–lime–PG and bentonite–lime–PG–Ss were investigated. On the basis of the results obtained, the following conclusions are reached:

- The hydration kinetics of the bentonite–lime and bentonite–lime–PG mixtures are slow.
- The formation of calcium silicate hydrate (C–S–H) increases the mechanical properties of bentonite–lime, bentonite–lime–PG and bentonite–lime–PG–Ss mixtures.
- The addition of PG alone to the bentonite–lime mixture allows the neo-formation of hydrated calcium silicate (C–S–H) and ettringite, which increases the compressive strength of the specimens.
- The microstructural analysis confirmed the formation of ettringite, which bonds the aggregates of the bentonite–lime–PG and bentonite–lime–PG–Ss mixtures.
- The addition of the black steel slag to the bentonite–lime–PG mixture accelerates the hydration kinetics and activates the pozzolanic reaction due to the presence of C₂S in the slag, and thus promoting the increase of the mechanical strength of the B–L–PG–Ss mixture.

- The compressive strength of hydrated BL, BL–PG and BL–PG–Ss mixtures increases from 15 to 28 days of curing. Furthermore, 8% of slag added to the B–L–PG mixture increase the mechanical compression of 0.6 MPa after 28 days of hardening.

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