



# (Micro)-structural comparison between geopolymers, alkali-activated slag cement and Portland cement

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## Abstract

Concurrently to research conducted on ordinary Portland cement (PC), new types of binders were developed during the last decades. These are formed by alkali-activation of metakaolin or ground-granulated blast furnace slag (GGBFS) and are named, respectively, geopolymers (GP) or alkali-activated slag (AAS). Four different cementitious materials were synthesised: PC, AAS, GP, and a mix GP-AAS and fully compared about their compositions and (micro)-structures.

X-ray diffraction has revealed the presence of semi-crystalline C–S–H gel binding phase in PC while AAS, GP and GP-AAS are nearly amorphous. Progressive structural changes have been observed between the different samples by means of infrared spectroscopy, <sup>29</sup>Si and <sup>27</sup>Al magic-angle-spinning nuclear magnetic resonance spectroscopy: there is a polymerisation extent of the (alumino)-silicate framework from PC [SiQ<sup>1</sup> and SiQ<sup>2</sup> units] to AAS [SiQ<sup>2</sup> and SiQ<sup>2</sup>(1Al) units] and finally to GP [SiQ<sup>4</sup>(2Al) and SiQ<sup>4</sup>(3Al) units]. Scanning electron microscopy has shown that GP is a homogeneous matrix while the other materials are composites containing GGBFS grains surrounded by a binding matrix. Energy dispersive X-ray EDX analyses (line scans) have shown the absence of formation of any specific phase at the matrix-grains interfaces.

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## 1. Introduction

The most used and most manufactured cement type is “Portland cement” (PC). It consists of the so-called Portland cement clinker mixed with a few percent of gypsum (CaSO<sub>4</sub>·2H<sub>2</sub>O) or anhydrite (CaSO<sub>4</sub>). Portland cement clinker contains four major phases: tricalcium silicate (C<sub>3</sub>S), β-dicalcium silicate (β-C<sub>2</sub>S), tricalcium aluminate (C<sub>3</sub>A), and ferrite solid solution (C<sub>2</sub>(A,F)) (in the cement nomenclature, C = CaO, S = SiO<sub>2</sub>, A = Al<sub>2</sub>O<sub>3</sub>, H = H<sub>2</sub>O, F = Fe<sub>2</sub>O<sub>3</sub>, M = MgO and  $\hat{S}$  = SO<sub>4</sub>). The main reaction product of Portland cement is a very poorly crystalline calcium silicate hydrate known as C–S–H. It is produced by hydration of C<sub>3</sub>S and β-C<sub>2</sub>S and is often designated as a gel when there are other phases admixed on a sub-micrometre scale. It is generally accepted that the Ca/Si molar ratio of C–S–H gel is in the range of 1.5–2, and in most cases near 1.7. C–S–H gel represents the primary binding phase in Portland cement and

controls the strength development of the paste. Great progress has been made in the understanding of the hydration chemistry of Portland cement.<sup>1</sup> However, at present, the structure of the C–S–H gel is imperfectly known and a lot of ambiguities still remain. What can be said unambiguously is that C–S–H has a layered structure with similarities to the 1.4-nm tobermorite [Ca<sub>5</sub>Si<sub>6</sub>O<sub>16</sub>(OH)<sub>2</sub>·8H<sub>2</sub>O or C<sub>5</sub>S<sub>6</sub>H<sub>9</sub>] and/or jennite [Ca<sub>9</sub>H<sub>2</sub>Si<sub>6</sub>O<sub>18</sub>(OH)<sub>8</sub>·6H<sub>2</sub>O or C<sub>9</sub>S<sub>6</sub>H<sub>11</sub>].<sup>11</sup> However, compared to 1.4-nm tobermorite and jennite, C–S–H is mainly amorphous, its C/S ratio is more variable and its silicate anion chains are very short (primarily dimers with a lesser amount of pentamers and octamers).<sup>12</sup>

Ground-granulated blast furnace slag (GGBFS) is an amorphous by-product of the steel industry. It has a latent hydraulic reactivity, which can be catalysed by proper activators such as Portland cement clinker, lime, gypsum, and alkali metal hydroxides, carbonates or silicates to form cementitious materials. The main reaction product generally cited for alkali-activated slag is C–S–H gel similar to that found in PC but with lower Ca/Si ratios (around 0.7).<sup>2–5</sup> Some authors have also mentioned the presence of a Mg–Al rich phase similar to hydrotalcite

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[Mg<sub>6</sub>Al<sub>2</sub>CO<sub>3</sub>(OH)<sub>16</sub>·4H<sub>2</sub>O] or a tetracalcium aluminate hydrate phase (C,M)<sub>4</sub>AH<sub>13</sub>.<sup>6,7</sup>

Alkali-activation of aluminosilicates, such as metakaolin or calcined clays, can also produce inorganic binders with excellent physical and chemical properties suitable for building purposes. These are called geopolymers and are widely developing since the last two decades. Aluminosilicates are reacted with aqueous alkaline solutions to activate and promote partial solid dissolution, which is then followed by polycondensation of the dissolved species to form an amorphous aluminosilicate gel and afterwards by the solid-state transformation of the gel. Within the geopolymer research community, the aluminosilicate gel is mainly considered as an amorphous three-dimensional framework of SiO<sub>4</sub> and AlO<sub>4</sub> tetrahedra randomly interlinked by shared O atoms. The negatively charged and tetrahedrally coordinated Al atoms inside the network are charge-balanced by alkali metal cations coming from the activating solution. Geopolymeric binders have thus similar chemical composition as natural zeolitic materials but without the extensive crystalline zeolitic structure.<sup>8–10</sup>

Ordinary Portland cement, alkali-activated slag cement, and geopolymers represent three different cementitious materials with their own characteristics and properties. These three types of materials have been extensively considered in the last 10 years. However, only specific studies about one type of cement have been considered. Up to now, no complete research has established precise relationships between geopolymers and the other two types of Ca-based binders.

The purpose of this study is a detailed comparison between geopolymers, alkali-activated slag cement and ordinary Portland cement regarding to the chemical composition, the crystal structure, and the microstructure. For the first time, differences and similarities between these three inorganic binders were highlighted by means of XRD analysis, IR spectroscopy, <sup>27</sup>Al and <sup>29</sup>Si MAS-NMR spectroscopy, and SEM-EDX analysis.

## 2. Experimental procedures

### 2.1. Reacting materials

Metakaolin, ground-granulated blast furnace slag (GGBFS), and Portland clinker were supplied by the Construction Steel Design Centre (ARCELOR INNOVATION, Liège) in Belgium. The chemical composition of both metakaolin and GGBFS is given in Table 1. Potassium metasilicate was purchased from PQ Europe (Maastricht, The Netherlands) and contains 11.1 wt.% of K<sub>2</sub>O, 23.9 wt.% of SiO<sub>2</sub>, and 65 wt.% of H<sub>2</sub>O. Deionised water and analytical grade KOH were used throughout all experiments.

Table 1  
Chemical composition of the starting materials

Components (wt.%)	Metakaolin	Ground-granulated blast furnace slag
SiO <sub>2</sub>	55.57	33–38
CaO		38–43
Al <sub>2</sub> O <sub>3</sub>	42.43	9–13
MgO		7–12
Fe		0.1–15
MnO		0–2
S		0–2.5
TiO <sub>2</sub>		0.3–0.9
K <sub>2</sub> O		0–0.8
H <sub>2</sub> O		<20

### 2.2. Synthesis

Geopolymer (GP) samples were synthesised by adding metakaolin to a solution of alkaline activators. This alkaline solution was prepared by the dissolution of potassium hydroxide pellets in a potassium metasilicate solution to obtain a molar oxide ratio K<sub>2</sub>O/SiO<sub>2</sub> equal to 0.68. A few millilitres of water were added at the same time to the paste to achieve a suitable workability. Alkali-activated slag (AAS) samples were produced by mixing GGBFS with the same alkaline solution (potassium metasilicate and potassium hydroxide, K<sub>2</sub>O/SiO<sub>2</sub> = 0.68) using a minimal amount of water (a few millilitres). The alkaline solution was also added to a mixture of metakaolin and GGBFS in order to prepare samples that contain both a geopolymeric phase and an alkali-activated slag phase (the GP-AAS samples). Portland cement (PC) samples were obtained by mixing Portland clinker CEM II (clinker + fly ash + GGBFS) with deionised water to achieve a water to cement ratio equal to 0.35.

After mixing for 5 min, all pastes were poured in hermetic polyethylene molds and allowed to cure at room temperature for more than 28 days. The amounts of the components used to synthesise the different samples (wt.%) are given in Table 2.

### 2.3. Experimental techniques

Infrared spectra were recorded on a Perkin-Elmer Fourier transform infrared (FTIR) 1720 × spectrometer using the KBr pellet technique (3.5 mg of powder sample mixed with 800 mg of KBr).

X-ray powder diffraction data were collected using a Siemens D5000 powder diffractometer (Karlsruhe, Germany) with Ni-filtered Cu K<sub>α</sub> radiation (40 kV; 50 mA).

<sup>29</sup>Si and <sup>27</sup>Al solid-state MAS-NMR spectroscopy was carried out at 9.04 T using a Bruker Avance DSX 400WB

Table 2  
Composition of the different samples: wt.% of the reactants

Samples	Metakaolin	GGBFS	Portland clinker CEM II	Alkaline solution K <sub>2</sub> O/SiO <sub>2</sub> = 0.68	H <sub>2</sub> O
GP	44.64			44.64	10.72
GP-AAS	25.24	37.87		25.27	11.63
AAS		66.51		28.50	4.99
PC			74.22		25.78

spectrometer (Karlsruhe, Germany) operating at 79.5 and 104.27 MHz for the  $^{29}\text{Si}$  and  $^{27}\text{Al}$  resonance frequencies, respectively.  $^{29}\text{Si}$  and  $^{27}\text{Al}$  MAS-NMR spectra (spinning rate of 10 kHz) were recorded with 7.8- and 3- $\mu\text{s}$  simple pulse and with 10- and 0.3-s delay, respectively. The number of transients collected is 1000 for  $^{27}\text{Al}$  and 1200 for  $^{29}\text{Si}$ . Chemical shifts are referenced to tetramethylsilane for the  $^{29}\text{Si}$  nuclei and to  $\text{AlCl}_3(1\text{M})$  for the  $^{27}\text{Al}$  nuclei.

Scanning electron microscopy (SEM) and energy dispersive X-ray (EDX) micro-analysis were performed on a Philips ESEM XL30 FEG (Eindhoven, The Netherlands) with an accelerating voltage of 10–15 kV. All cementitious samples were enclosed in epoxy resin and polished with SiC papers and diamond paste; the polished cross-sections were then gold-coated before observation.

### 3. Results and discussions

#### 3.1. X-ray diffraction analysis

X-ray patterns of geopolymer sample (GP), geopolymer-alkali-activated slag sample (GP-AAS), alkali-activated slag sample (AAS), and Portland cement sample (PC) are presented in Fig. 1(A–D), respectively.

The PC sample is mainly amorphous in view of the large diffuse diffraction peak centred at  $30^\circ$  ( $2\theta$ ) (Fig. 1(D)).

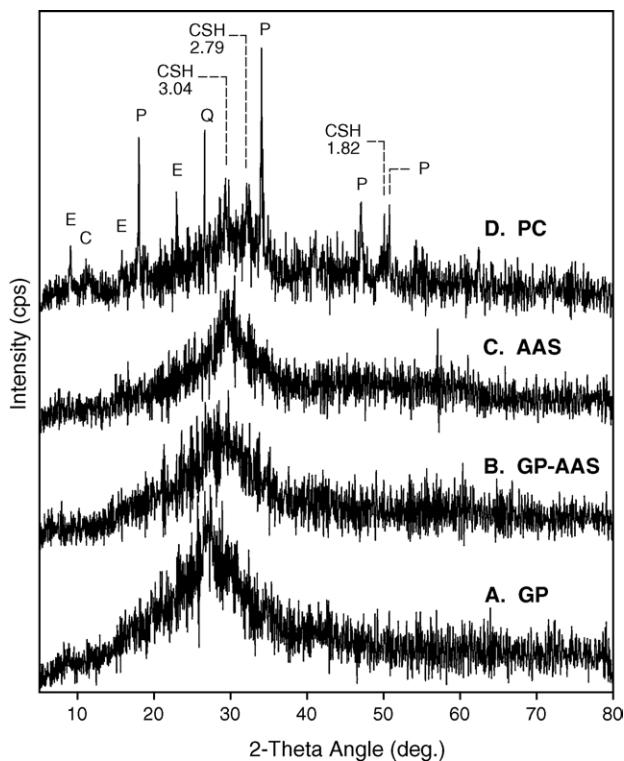


Fig. 1. X-ray powder patterns of (A) GP sample, (B) GP-AAS sample, (C) AAS sample, and (D) PC sample [E = ettringite ( $\text{Ca}_6\text{Al}_2(\text{SO}_4)_3(\text{OH})_{12}\cdot 26\text{H}_2\text{O}$ ), PDF no. 41–1451; P = portlandite ( $\text{Ca}(\text{OH})_2$ ), PDF no. 44–1481; CSH = calcium silicate hydrate ( $1.5\text{CaO}\cdot\text{SiO}_2\cdot x\text{H}_2\text{O}$ ), PDF no. 33–0306; Q = quartz ( $\text{SiO}_2$ ), PDF no. 85–1780; C = calcium aluminate hydrate ( $\text{Ca}_4\text{Al}_2\text{O}_7\cdot 13\text{H}_2\text{O}$  or  $\text{C}_4\text{AH}_{13}$ ), PDF no. 11–0203].

However, for the PC sample only, the X-ray pattern also shows some sharper reflection peaks corresponding to ettringite [ $\text{Ca}_6\text{Al}_2(\text{SO}_4)_3(\text{OH})_{12}\cdot 26\text{H}_2\text{O}$ ], portlandite [ $\text{Ca}(\text{OH})_2$ ], quartz ( $\text{SiO}_2$ ), and also to calcium silicate hydrate gel [ $(\text{CaO})_x\cdot\text{SiO}_2\cdot(\text{H}_2\text{O})_y$ ], respectively. C–S–H gel is characterised by three diffraction peaks at 3.04, 2.79, and 1.82 Å (representing the (2 2 0), (4 0 0), and (0 4 0) reflections of 1.1-nm tobermorite, respectively), which are reasonably well visible on the X-ray pattern of PC sample (Fig. 1(D)).<sup>1,13</sup>

As shown by the diffuse halo diffraction peak centred around  $30^\circ$  ( $2\theta$ ), GP, GP-AAS, and AAS samples (Fig. 1(A–C), respectively) are completely amorphous with clearly a lack of long-range order. The halo undergoes progressively a slight shift towards greater angle values from about  $27^\circ$  ( $2\theta$ ) for GP sample to about  $31^\circ$  ( $2\theta$ ) for PC sample.

Although many authors have found that the major hydration products of AAS are C–S–H gel phases,<sup>3,4</sup> X-ray patterns of GP-AAS and AAS samples do not allow such a conclusion. Moreover, there is no trace of a Mg,Al rich phase related to the naturally occurring mineral hydrotoalcite [ $\text{Mg}_6\text{Al}_2(\text{CO}_3)(\text{OH})_{16}\cdot 4\text{H}_2\text{O}$ ] and of a tetracalcium aluminate hydrate phase such as  $(\text{C},\text{M})_4\text{AH}_{13}$ , despite these are also referred as current hydration products of AAS.<sup>6</sup>

On the other hand, the X-ray pattern of GP sample is very similar to that previously published for geopolymer materials.<sup>8–10</sup>

X-ray diffraction allows only to say that absolutely no crystalline or semi-crystalline phase is formed in GP, GP-AAS, and AAS materials, which are more amorphous than the C–S–H gel phase occurring in ordinary Portland cement, with a lack of long-range order.

#### 3.2. Infrared analysis

Fig. 2(A–D) display infrared spectra of geopolymer sample (GP), geopolymer-alkali-activated slag sample (GP-AAS), alkali-activated slag sample (AAS), and Portland cement sample (PC), respectively. These spectra are very comparable to those previously published for each type of sample.<sup>2,8,15</sup>

Infrared spectra of the four samples are rather similar, presenting analogous absorption bands. All show bands at  $3440$  and  $1650\text{ cm}^{-1}$ , respectively, related to O–H stretching and bending modes of molecular water; and also near  $1000\text{ cm}^{-1}$  and at  $450\text{ cm}^{-1}$  due to anti-symmetric Si–O(Al) stretching vibrations ( $\nu_3$ ) and to in-plane Si–O bending vibrations ( $\nu_2$ ) in  $\text{SiO}_4$  tetrahedra, respectively. Nevertheless, some differences are observed. Only the infrared spectrum of the PC sample (Fig. 2(D)) displays a sharp band at  $3640\text{ cm}^{-1}$  associated to O–H stretching vibrations of portlandite [ $\text{Ca}(\text{OH})_2$ ]; a weak band at  $1110\text{ cm}^{-1}$  and a sharp peak at  $660\text{ cm}^{-1}$  corresponding, respectively, to S–O<sub>4</sub> stretching mode ( $\nu_3$ ) and out-of-plane bending mode ( $\nu_4$ ) of ettringite; and also a shoulder near  $530\text{ cm}^{-1}$  probably due to out-of-plane Si–O bending vibrations ( $\nu_4$ ).<sup>16,17</sup>

All samples with the exception of the GP sample contain carbonate species pointed out by the presence of the large absorption band near  $1450\text{ cm}^{-1}$  and of the small one at  $870\text{ cm}^{-1}$ ,

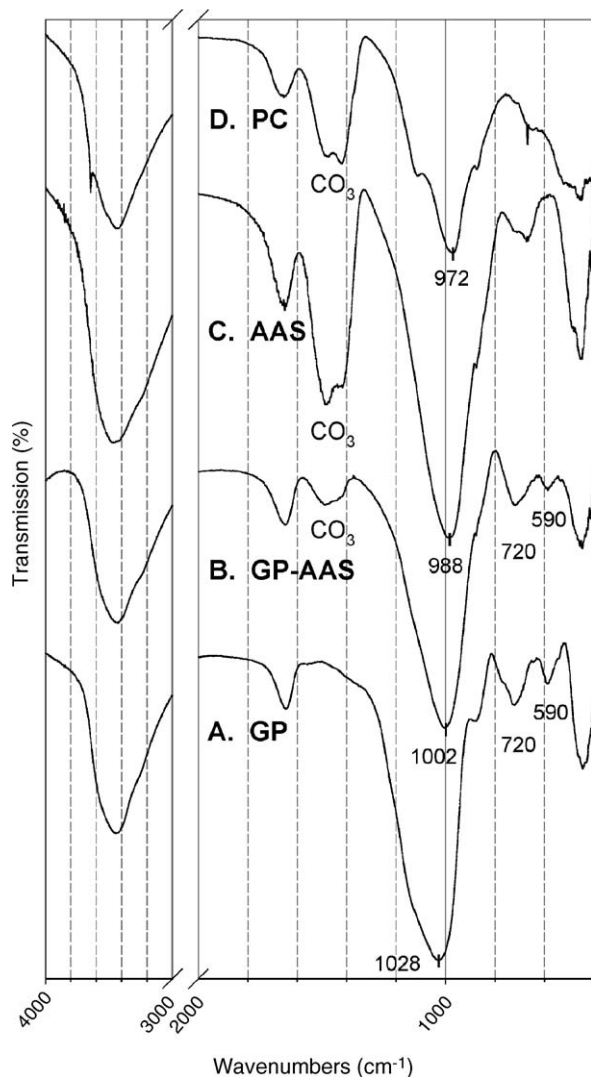


Fig. 2. Infrared spectra of (A) GP sample, (B) GP-AAS sample, (C) AAS sample, and (D) PC sample.

respectively, related to anti-symmetric stretching ( $\nu_3$ ) and out-of-plane bending ( $\nu_2$ ) modes of  $\text{CO}_3^{2-}$  ions.<sup>15,16</sup> This observation suggests that geopolymers are less sensitive to atmospheric carbonation than ordinary Portland cement or alkali-activated slag cement.

The principal band associated with the Si–O(Al) stretching vibrations in  $\text{SiO}_4$  tetrahedra near  $1000\text{ cm}^{-1}$  is very broad particularly for the GP, GP-AAS, and AAS samples. It confirms that these materials are more disordered than Portland cement. The lack of a sharp feature is indicative of the wide distribution of the  $\text{SiQ}^n(\text{mAl})$  units occurring in these structures. The Si–O stretching modes for the  $\text{SiQ}^n$  units show infrared absorption bands localised around 1200, 1100, 950, 900, and  $850\text{ cm}^{-1}$  for  $n=4, 3, 2, 1,$  and  $0$ , respectively.<sup>18</sup> These values shift to lower wavenumbers when the degree of silicon substitution by aluminium in the second coordination sphere increases, as a consequence of the weaker Al–O bonds.

In Fig. 2, it appears that this Si–O stretching band shifts progressively towards greater wavenumbers from  $970\text{ cm}^{-1}$  for

PC sample (Fig. 2(D)) to  $1030\text{ cm}^{-1}$  for GP sample (Fig. 2(A)) as the Si/Al molar ratio is decreasing (4 to 1.8). These results indicate a distribution of the  $\text{Q}^n$  units centred around  $\text{Q}^1$  and  $\text{Q}^2$  units for Portland cement sample, while the shift to higher wavenumbers points out the presence of more polymerised units such as  $\text{Q}^2$  and  $\text{Q}^3$  units for alkali-activated slag sample and  $\text{Q}^4$  units for geopolymer sample.

In the structure of highly polymerised silicates, silicoxygen rings occur most frequently with three, four and six members. The typical bands of the ring system (pseudo-lattice vibrations) are located mainly in the middle IR spectra in the range  $800\text{--}500\text{ cm}^{-1}$ :  $720\text{--}700\text{ cm}^{-1}$  for 3-membered rings,  $650\text{ cm}^{-1}$  for four-membered rings and  $620\text{--}600\text{ cm}^{-1}$  in the case of six-membered rings.<sup>19,20</sup>

Two absorption bands at  $720$  and  $590\text{ cm}^{-1}$  appear in the case of GP and GP-AAS samples (Fig. 2(A and B)). The higher band would be related to the symmetric stretching vibrations of the Si–O–Si(Al) bridges. The next band at  $590\text{ cm}^{-1}$  would correspond to the cyclosilicates vibrations; its intensity increases with the aluminium and silicon contents (Table 3). The analogy between the IR spectra of the GP sample and  $\alpha$ -cristobalite  $\text{SiO}_2$  (the low temperature tetragonal phase) allows to suggest a structural similitude with interconnected cyclohexasilico-aluminate rings.<sup>21</sup> The width of this  $590\text{ cm}^{-1}$  band can be related to disorder and/or deformation of the six-membered rings.

In the AAS spectrum, the absorption band around  $700\text{ cm}^{-1}$  suggests the formation of smaller three-membered rings and a less polymerised structure.<sup>20</sup>

Finally, GP sample and GP-AAS sample show very similar infrared spectra; GP-AAS spectrum marks a transition between GP spectrum and AAS spectrum. The spectrum of AAS sample is slightly different with an intermediate shape between PC sample and GP sample. It suggests that the structure of alkali-activated slag phases is a transition between the structures of ordinary Portland cement and

Table 3

Composition of the matrix in GP, GP-AAS, AAS, and PC samples; and of the GGBFS particles in GP-AAS, AAS, and PC samples, detected by SEM/EDX

Elements (oxides) (wt.%)	GP sample	GP-AAS sample	AAS sample	PC sample
<b>Matrix</b>				
MgO		3.5	4	5
Al <sub>2</sub> O <sub>3</sub>	27	22	8.8	7
SiO <sub>2</sub>	58	49.8	43	33
K <sub>2</sub> O	14	8.9	14.4	1.6
CaO		16	29.2	53
Ca/Si (mol)	–	0.3	0.7	1.7
Si/Al (mol)	1.8	1.9	4.1	4.0
<b>Particles</b>				
MgO		7.19	7.3	8
Al <sub>2</sub> O <sub>3</sub>		12.6	11.9	11.8
SiO <sub>2</sub>		39.6	39	37.5
K <sub>2</sub> O		0.9	1.09	0.51
CaO		40	40.7	42
Ca/Si (mol)		1.1	1.1	1.2
Si/Al (mol)		2.7	2.8	2.7

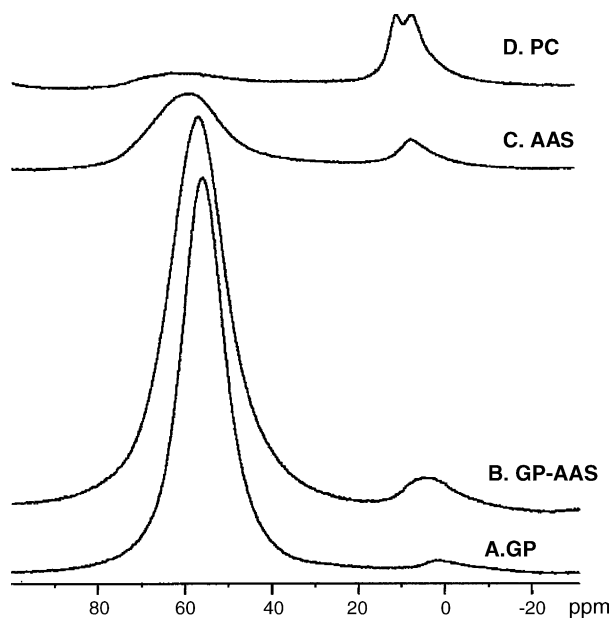


Fig. 3.  $^{27}\text{Al}$  MAS-NMR spectra of (A) GP sample, (B) GP-AAS sample, (C) AAS sample, and (D) PC sample.

geopolymers while the incorporation of a non-negligible quantity of GGBFS (around 40 wt.%) in a geopolymeric binder seems to have no fundamental influence on the final structure of geopolymers with regard to the similarity between the infrared spectra (above  $400\text{ cm}^{-1}$ ) of GP sample and GP-AAS sample.

### 3.3. $^{29}\text{Si}$ and $^{27}\text{Al}$ MAS-NMR analysis

$^{29}\text{Si}$  and  $^{27}\text{Al}$  MAS-NMR appear as suitable techniques to probe structural features of geopolymers, alkali-activated slag cement, and Portland cement. Actually, determination of the aluminium coordination number with oxygen (four, five, or six) is possible with  $^{27}\text{Al}$  MAS-NMR<sup>22</sup> and  $^{29}\text{Si}$  MAS-NMR allows the differentiation between the different  $\text{SiQ}^n(\text{mAl})$  structural units.<sup>23</sup>

$^{27}\text{Al}$  MAS-NMR spectra of GP, GP-AAS, AAS, and PC samples are shown on Fig. 3(A–D), respectively.

In the  $^{27}\text{Al}$  MAS-NMR spectrum of PC sample (Fig. 3(D)), the very large and weak band centred approximately at 64 ppm is assigned to tetrahedrally coordinated Al incorporated in the C–S–H gel, while the spectral range for octahedrally coordinated Al displays two sharper bands at 12 and 8 ppm. The first sharp resonance originates from ettringite ( $\text{C}_6\text{A}\hat{\text{S}}_3\text{H}_{32}$ ) (reported to have an isotropic chemical shift of 13.1 ppm) that has already been detected by X-ray diffraction (Fig. 1(D)). The second one at 8 ppm could correspond to the isotropic chemical shift of the monosulfate phase,  $\text{C}_4\text{A}\hat{\text{S}}\text{H}_{12}$  ( $\delta_{\text{iso}} = 11.8$  ppm), and of the tetracalcium aluminate hydrate phase,  $\text{C}_4\text{AH}_{13}$  ( $\delta_{\text{iso}} = 10.2$  ppm) (the chemical shift difference can be due to a large amount of disorder as these two phases are unobserved by XRD).<sup>24,25</sup>

$^{27}\text{Al}$  MAS-NMR spectrum of AAS sample is very different (Fig. 3(C)); it shows a major content of tetrahedrally coordinated

Al in connection with the large band centred at 59 ppm and a lower amount of octahedrally coordinated Al related to the band at 8 ppm. The later resonance is possibly assigned to a tetracalcium aluminate hydrate phase [ $(\text{C},\text{M})_4\text{AH}_{13}$ ] or to a Mg–Al rich phase similar to hydrotalcite [ $\text{Mg}_6\text{Al}_2\text{CO}_3(\text{OH})_{16}\cdot 4\text{H}_2\text{O}$ ], often referred as reaction products of alkali-activated slag.<sup>7</sup> This phase, substantially disordered, would be formed consequently to the alkaline activation of GGBFS, as  $^{27}\text{Al}$  MAS-NMR spectrum of raw GGBFS displays only a large resonance near 57 ppm.

GP and GP-AAS samples have very similar  $^{27}\text{Al}$  MAS-NMR spectra. Both show a well-resolved resonance at 56–57 ppm due to Al in four-fold coordination and a light hump in the base line around 2–5 ppm, especially for GP-AAS sample. This later is associated to Al in octahedral coordination and is certainly unreacted aluminium from metakaolin.<sup>8</sup> The bands associated to Al in four-fold coordination have very close chemical shifts (56–57 ppm) to that of AAS sample (59 ppm), which shows similar chemical environments of Al(IV) in GP, GP-AAS, and AAS samples.

The comparison between  $^{27}\text{Al}$  MAS-NMR spectra of the four samples highlights three general continuous evolutions from Portland cement to geopolymers:

1. An increase of the total resonance intensities consistent with a corresponding raise of the samples Al content.
2. An increase of the tetrahedrally coordinated Al content and a corresponding decrease for Al in six-fold coordination.
3. An increase in the polymerisation degree of the aluminosilicate framework as the maximum of the Al(IV) peak moves from 64 ppm for PC sample ( $\text{AlQ}^2(2\text{Si})$  or  $\text{AlQ}^3(3\text{Si})$  structural units) to 56 ppm for GP sample ( $\text{AlQ}^4(4\text{Si})$  structural unit).

$^{29}\text{Si}$  MAS-NMR spectra of GP, GP-AAS, AAS, and PC samples are shown on Fig. 4(A–D), respectively.

Three major environments can be deduced from the  $^{29}\text{Si}$  MAS-NMR spectrum of PC sample. The weaker is characterised by a large band centred approximately at  $-73$  ppm arising from raw Portland cement clinker. The two other main resonances appear at approximately  $-80$  and  $-86$  ppm and originate from the  $\text{SiQ}^1$  (disilicates or chain end groups) and  $\text{SiQ}^2$  (middle groups in chains) units, respectively.<sup>24</sup>

$^{29}\text{Si}$  MAS-NMR spectrum of AAS sample (Fig. 4(C)) shows some non-negligible differences compared to that of PC sample. It displays only two main peaks, the most important resonates at  $-86$  ppm and the second one appears at  $-83$  ppm. These have been previously assigned, respectively, to  $\text{SiQ}^2$  and  $\text{SiQ}^2(1\text{Al})$  units by a lot of authors who published similar spectra for alkali-activated slags.<sup>2,3,5,7,26</sup> So, most of the silicate tetrahedra are chain mid-members and important Al substitution for Si occurs in the silicate chains. The general characteristic of the AAS sample spectrum compared to the PC sample is a shift towards more negative chemical shifts, which reflects significantly longer aluminosilicate chains than in Portland cement.

GP-AAS sample has a slightly different  $^{29}\text{Si}$  MAS-NMR spectrum (Fig. 4(B)) compared to those of AAS and PC samples.

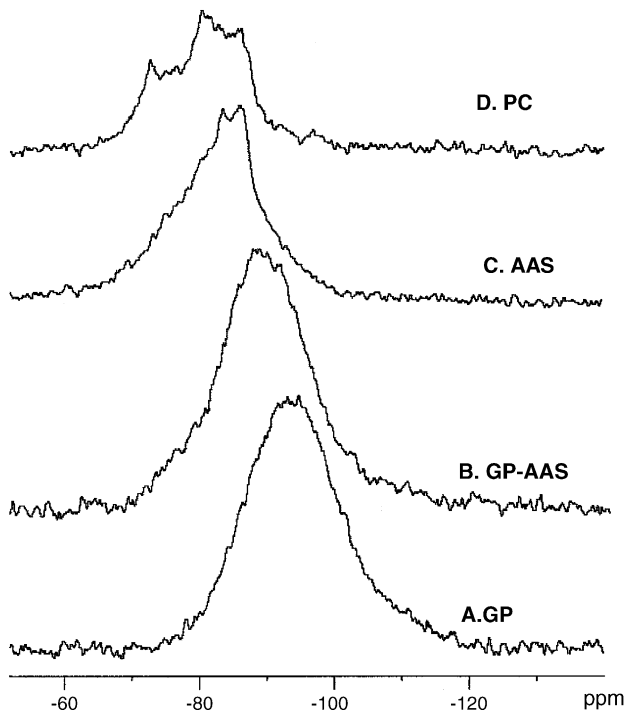


Fig. 4.  $^{29}\text{Si}$  MAS-NMR spectra of (A) GP sample, (B) GP-AAS sample, (C) AAS sample, and (D) PC sample.

No distinct peak can be distinguished but only a single broad resonance characteristic of amorphous material containing a range of slightly different environments. The peak is centred approximately at  $-89$  ppm, consistent with chain branching sites ( $\text{SiQ}^3$

units) and three-dimensional cross-linked sites ( $\text{SiQ}^4$  units), both with substantial Al substitution for Si.<sup>8</sup>

$^{29}\text{Si}$  MAS-NMR spectrum of GP sample is very similar to that of GP-AAS sample (Fig. 4(A)). The only difference lies in the position of the peaks, whose maxima are localised at  $-89$  ppm for GP-AAS sample and at  $-95$  ppm for GP sample. Geopolymer materials are thus characterised by a highly polymerised structure composed principally of three-dimensional cross-linked sites such as  $\text{SiQ}^4(2\text{Al})$  and  $\text{SiQ}^4(3\text{Al})$  units.<sup>9,10</sup>

In summary, Portland cement is composed mainly of  $\text{SiQ}^1$  and  $\text{SiQ}^2$  species, which is consistent with the models proposed previously for the C–S–H gel. In these models, silicate groups organize in linear finite chains of “dreierketten” structure (principally dimers and also some pentamers and octamers). In alkali-activated slag, C–S–H gel is formed by longer chains whereas  $^{29}\text{Si}$  MAS-NMR spectrum shows mainly chain mid-member sites [ $\text{SiQ}^2$  and  $\text{SiQ}^2(1\text{Al})$ ]. On the contrary, geopolymers contain essentially three-dimensional cross-linked sites [ $\text{SiQ}^4(2\text{Al})$  and  $\text{SiQ}^4(3\text{Al})$ ], which reflects a highly polymerised structural aluminosilicate framework. Comparison between  $^{29}\text{Si}$  MAS-NMR spectra of GP and GP-AAS samples shows that the incorporation of GGBFS in a geopolymer phase does not cause any significant structural change except a slight framework depolymerisation.

The above results are in complete agreement with the conclusions obtained from the infrared and  $^{27}\text{Al}$  MAS-NMR spectroscopies; a continuous increase of the degree of polymerisation from Portland cement to alkali-activated slag and to geopolymers.

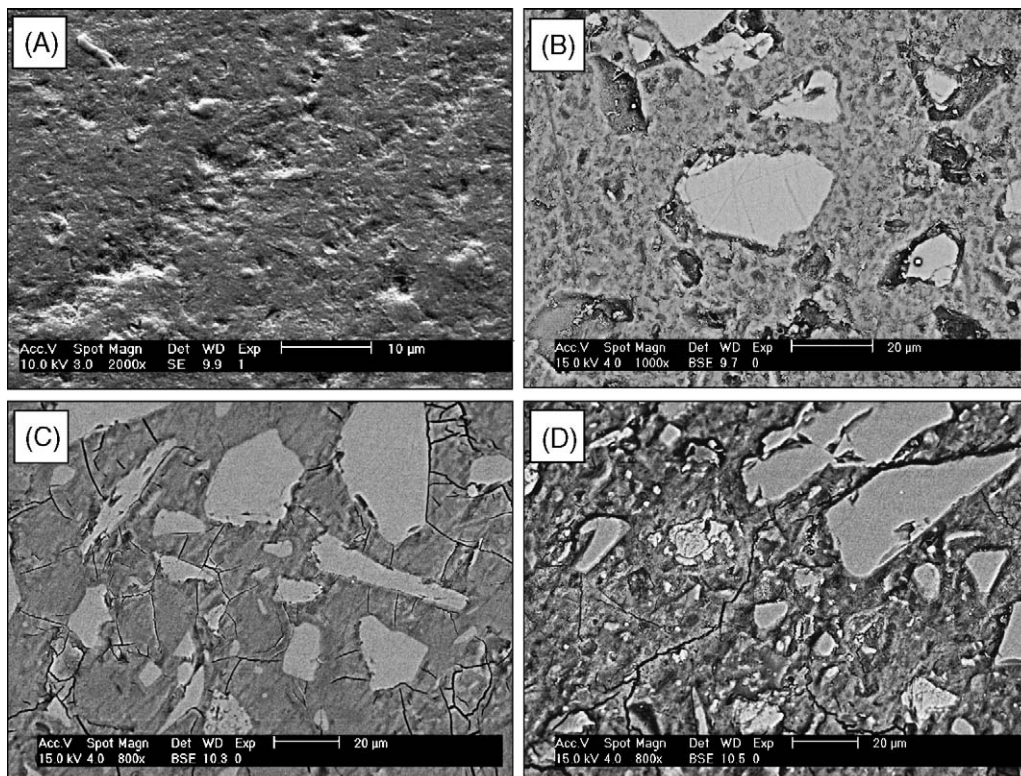


Fig. 5. SEM image of GP sample (A), GP-AAS sample (B), AAS sample (C), and PC sample (D).

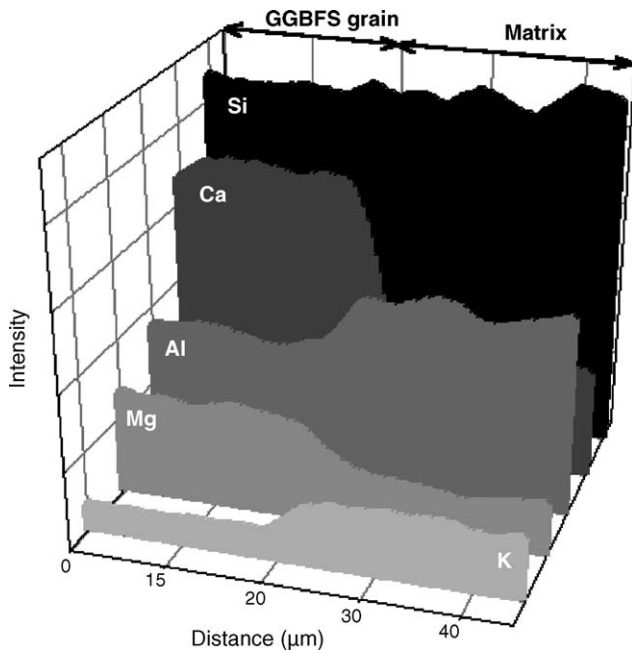


Fig. 6. SEM/EDX line scans from a GGBFS grain to the binding matrix in GP-AAS sample.

3.4. SEM analysis

The micrographs of GP, GP-AAS, AAS, and PC samples are shown on Fig. 5(A–D), respectively.

GP micrograph displays a very homogeneous binding matrix without aggregate. On the contrary, GP-AAS, AAS, and PC samples are more heterogeneous. Their micrographs show particles of various shapes and sizes surrounded by a cementitious matrix. GP and GP-AAS materials show absolutely no crack, in opposi-

tion to PC sample and especially to AAS sample. The cracks are running in the binding matrix between the slag grains, but they never pass through them and they are generally not located at the grain-matrix interface either. No reaction zone is observed at the interface between the grains and the matrix around the particles. The cementitious matrix in GP-AAS and AAS samples displays a vitreous aspect, more homogeneous than in PC sample.

Energy dispersive X-ray (EDX) analyses have been performed to determine the composition of the matrix and of the particles. The results are listed in Table 3, where each value is the mean of five measurements realised either in the matrix (top of the Table 3) or in the particles (bottom of the Table 3). In view of their composition, all the particles are GGBFS grains. GP matrix contains only Si, Al, and K arising from the dissolution of metakaolin in the presence of KOH and sodium metasilicate. Consequently to the incorporation of GGBFS, GP-AAS matrix contains also Mg and Ca that may only result from the partial dissolution of GGBFS particles. This confirms previous results suggesting that smaller GGBFS grains react totally while the larger ones react partially with alkaline activators and contribute to the formation of the binding matrix of the composite.<sup>8</sup> In AAS sample, GGBFS also dissolves partially and contributes to the binding phase in view of the presence of Mg, Al, and Ca in the AAS matrix. PC matrix contains essentially Ca and Si with a Ca/Si molar ratio equal to 1.7, and also lower quantities of Al, Mg, and K. Comparatively, AAS and GP-AAS matrix have much lower Ca/Si molar ratios, respectively, equal to 0.7 and 0.3.

Line scans (SEM-EDX) have been performed from a GGBFS particle to the binding matrix for GP-AAS, AAS, and PC samples, with the purpose to study the evolution of Si, Ca, Al, Mg, and K concentrations through the grain-matrix interface. These data are shown, respectively, in Figs. 6–8. The amount of K is

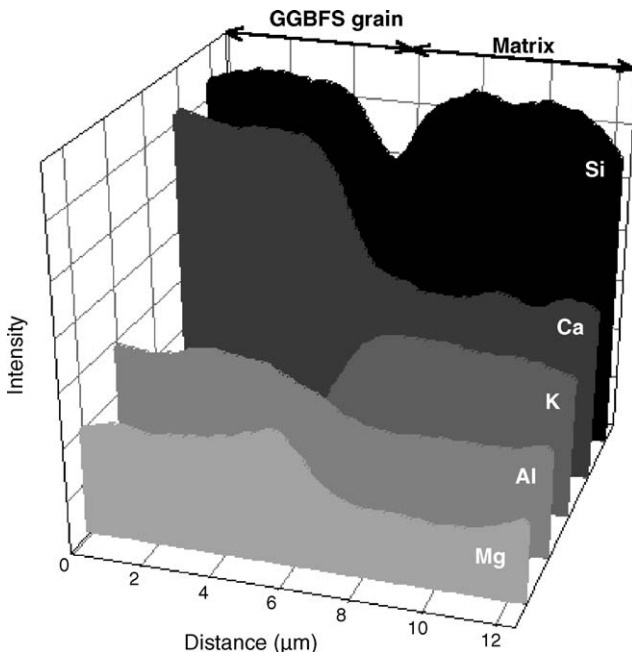


Fig. 7. SEM/EDX line scans from a GGBFS grain to the binding matrix in AAS sample.

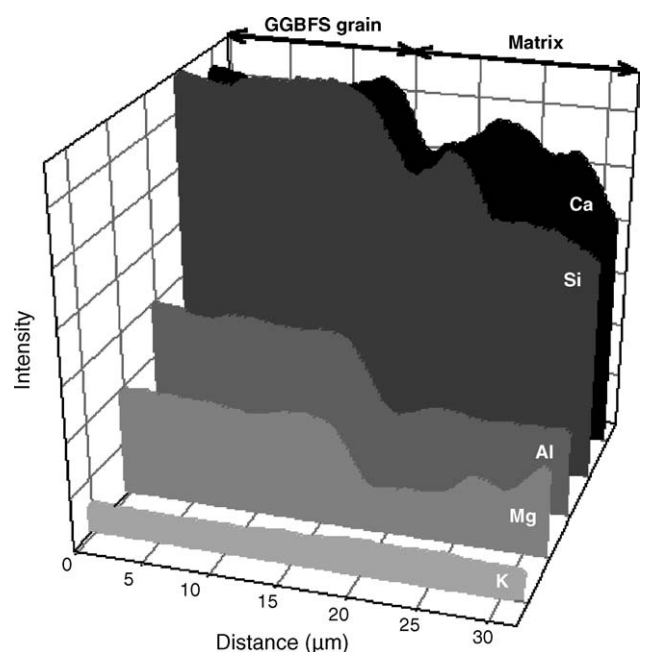


Fig. 8. SEM/EDX line scans from a GGBFS grain to the binding matrix in PC sample.

very low in all the three GGBFS grains; it increases a little in the GP-AAS matrix and much more in the AAS matrix; on the other hand, it remains constant in the PC matrix in relation to the alkalinity of the pastes: AAS > GP-AAS > PC. For GP-AAS and AAS samples, Mg and Ca decrease progressively from the grain to the matrix where they are found in non-negligible quantities, as a result of slag dissolution. In PC line scan, all the elements concentrations (except for K) diminish from slag grain to the matrix where some fluctuations are detectable, probably due to the heterogeneous character of the PC matrix.

In any of the three materials, no distinct phase with a particular composition is formed on the slag particles surface. No specific compound can be identified at the GGBFS grain-matrix interface by means of EDX line scan.

#### 4. Conclusions

Four different types of cementitious materials have been synthesised in order to perform a complete and detailed comparison about their composition, crystallographic structure, and microstructure. These are an ordinary Portland cement (PC) sample, an alkali-activated slag (AAS) sample, a geopolymer (GP) sample, and finally a sample containing both alkali-activated slag phases and geopolymer phases (GP-AAS). For the first time, these different inorganic binders have been extensively studied to highlight differences and similarities in their (micro)-structure.

X-ray diffraction has shown that PC sample contains C–S–H gel phase, ettringite, portlandite, and quartz. On the contrary, absolutely no crystalline or semi-crystalline phase is formed in GP, GP-AAS, and AAS samples. These are more amorphous than the C–S–H gel phase in Portland cement.

The four cementitious materials show analogous absorption bands in their infrared spectra even if differences are observed. Only the PC spectrum displays sharp bands corresponding to portlandite and ettringite. All samples except for GP sample contain carbonate species, which points out a greater resistance of geopolymers to carbonation. From PC sample to GP sample, the large Si–O stretching band (near  $1000\text{ cm}^{-1}$ ) shifts progressively to greater wavenumbers, which indicates an increase of the  $\text{SiO}_4$  tetrahedra polymerisation. GP and GP-AAS samples contain most likely six-membered alumino-silicate rings in their structure while AAS sample is composed of three-membered rings. On the contrary, no ring was observed in the structure of PC sample.

$^{27}\text{Al}$  MAS-NMR spectroscopy has shown progressive structural changes from PC sample to GP sample. There is an increase of the Al(IV) content and a subsequent decrease for Al(VI); and especially a shift of the Al(IV) peak to lower chemical shifts, which indicates a polymerisation extent of the (alumino)-silicate framework (from  $\text{AlQ}^2(2\text{Si})$  to  $\text{AlQ}^4(4\text{Si})$ ).  $^{27}\text{Al}$  spectrum of PC sample has also confirmed the presence of ettringite and the incorporation of Al(IV) in C–S–H gel. The spectrum of AAS sample has detected the presence of Al(VI) arising probably from a hydration product of AAS such as  $(\text{C},\text{M})_4\text{AH}_{13}$  or a Mg–Al rich phase, similar to hydrotalcite.

$^{29}\text{Si}$  MAS-NMR spectroscopy has confirmed previous results obtained by IR and  $^{27}\text{Al}$  MAS-NMR spectroscopies: there is a progressive polymerisation extent of the silicate network from Portland cement to alkali-activated slag cement, and further to geopolymers. In Portland cement, C–S–H gel contains silicate groups organised in linear finite chains of “dreierketten” structure, thus mainly  $\text{SiQ}^1$  and  $\text{SiQ}^2$  species. In alkali-activated slag cement, C–S–H gel is formed by longer chains in view of the predominance of chain mid-member units,  $\text{SiQ}^2$  and  $\text{SiQ}^2(1\text{Al})$ . Geopolymer materials are characterised by a highly polymerised aluminosilicate structure composed mainly of three-dimensional cross-linked units,  $\text{SiQ}^4(2\text{Al})$  and  $\text{SiQ}^4(3\text{Al})$ .

The elemental compositions and the microstructures of the four types of cement have been analysed by means of SEM-EDX. Geopolymer sample appears as a homogeneous material while the other samples are composite materials composed of GGBFS particles surrounded by a binding matrix. Absolutely no crack is observed in geopolymer materials contrary to alkali-activated slag and to Portland cement. Portland cement matrix appears more heterogeneous than the three others, the former having a vitreous aspect and a very constant chemical composition. AAS matrix and GP-AAS matrix have lower Ca/Si molar ratios (0.7 and 0.3, respectively) than PC matrix (1.7), which is consistent with their longer or more polymerised (alumino)-silicate chains. Line scans–EDX analyses have shown the absence of any distinct phase with a specific chemical composition at the matrix-particle interfaces and have confirmed the partial dissolution of the slag grains, involved in the formation of the binding phase.

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